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6. a) NH_4^+ + b) OH^- - c) A conjugate acid-base pair are two chemical species that differ by a proton. ChemActivity 42 Answers to Problems and Exercises 1. SO_4^{2-} ; CO_3^{2-} ; OH^- ; O^{2-} ; H_2O ; NH_3 ; CH_3NH_2 ; F^- ; CH_3COO^- . 2. HSO_4^- ; HCO_3^- ; H_3O^+ ; H_2O ; OH^- ; NH_4^+ ; CH_3NH_3^+ ; HCN ; CH_3COOH ; HF ; H_2CO_3 ; NH_3 .

ANSWERS_to_CTQs_and_Exercises_Chemactivities_42_43

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4. The bond energy (in MJ/mole) is the energy required to break one mole of the specified bonds. Triple bonds share six electrons

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between two atoms and are stronger (have greater bond energy) than double bonds, which share four electrons between two atoms.

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PLEASE NOTE: If you have a question about these answers, it is your responsibility to come to office hours or ask during class work time. UNIT 12 - HW Practice Keys - ChemActivity 51: Cell Voltage - ChemActivity 50: Electrochemical Cell UNIT 11 - HW Practice Keys - ChemQuest 55: Free Energy - ChemQuest 54: 2nd Law of Thermodynamics

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Types of Chemical Reactions Do atoms rearrange in predictable patterns during chemical reactions? why? Recognizing patterns allows us to predict future behavior.

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patterns to predict dangerous storms so people can get their families to safety. Political analysts use patterns to predict election outcomes. Similarly, chemists classify chemical equations according to their ...

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4. The answer here depends on the students' answer to CTQ 3.

5.a) The IE 1 for He is greater than the IE 1 for H because the nuclear charge on a He atom is +2 whereas the nuclear charge on H is +1. b) The IE 1 for Li is less than the IE 1 for He because at least one of the electrons of Li must be farther away from the nucleus than any electron ...

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ChemActivity 42 Know the Bronsted-Lowery definition of an acid and a base. In a chemical reaction identify the Bronsted acid and base in the reactants. Determine the conjugate base of a given acid and determine the conjugate acid of a base.

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For example: $94\text{Be} + 42\text{He} + 14\text{N} + 42\text{He} + 1\text{H} + 21\text{Sc} + 42\text{He}$ In nuclear equations, the total number of positive charges (represented by the atomic

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numbers) of the reactants (substances on the left of the reaction arrow) equals the total number of positive charges of the products (substances on the right of the reaction arrow).

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