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## **Electrochemical Cells Ap Chem Lab**

The lab is done in three parts. In Part 1, a table listing the reduction potentials of metal ions is made. In part 2, the Nerst equation is used to measure the voltage

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Introduction Oxidation-  
reduction reactions  
form a major class of  
chemical reactions.  
From the reactions of  
oxygen with sugars,  
fats, and proteins that

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provide energy for life to the corrosion of metals, many important reactions involve the processes of oxidation and reduction.

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I. Oxidation-reduction  
reactions form a major  
class of chemical  
reactions. From the  
reactions of oxygen  
with sugars, fats, and  
proteins that provide  
energy for life to the  
corrosion of metals,  
many

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views.

## **Electrochemical Cell Lab**

Before you begin, save  
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Template on your  
computer as

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Title: Electrochemical  
Cells.

Purpose/Hypothesis: To



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understand the function of electrochemical cells. To recognize the relation between reduction and oxidation reactions. To determine the relative reduction potential of sample metals. To calculate reduction potentials

## **Electrochemistry**

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Whataffects(theamoun  
tofmetalplating(outont  
he

**Name! ! ! ! !**

**Lab!Section: ! Electr  
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!!!!**

One can determine the  
standard potential of  
any electrochemical  
cell by: 1. Identifying  
the oxidation (anode)  
and reduction

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(cathode) half-cells. 2.

Looking up the  
standard half-cell  
potentials in a table of  
reduction potentials.  
An abbreviated table is  
included at the end of  
this lab procedure.

## **Lab 10 - Electrochemical Cells**

AP Chemistry learning  
Objectives. 3.12 The  
student can make  
qualitative or  
quantitative

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predictions about galvanic cell reactions based on half-cell reactions and electrode potentials. 3.13 The student can analyze data regarding galvanic cells to identify properties of the underlying REDOX reactions.

## **Electrochemical Cell Demonstration**

### **Voltaic Cell: Zinc ...**

Sketch how the  $Zn^{2+}$

$(aq)/Cu(s)$

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electrochemical cell in Model 1 may appear in a lab setup. Label the electrodes and solutions. Include a voltmeter in your drawing.  $\text{Zn(s)}$   $\text{Zn}^{2+}(\text{aq})$  1.100 v  $\text{Cu(s)}$   $\text{Cu}^{2+}(\text{aq})$  5. Is the reaction in Model 1 at equilibrium at any point during the experiment?

## **Hooper's Laboratory - Home**

E° cell, using a Vernier

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voltage probe as shown in Figure 3. You will use 1.0 M solutions for both half-cells, so  $Q = 1$  and  $\ln Q = 0$  for the reaction. Thus the cell potential measured will be the same as  $E^\circ_{\text{cell}}$  as evident from the Nernst equation (6). You will then use your UCCS Chem 106 Laboratory Manual Experiment 9

## **Experiment 9** **Electrochemistry I -**

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## **Galvanic Cell**

Middle East Technical  
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OpenCourseWare [  
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] Chemistry

Department 12.  
Electrochemistry -  
Voltaic Cells Course  
Link: <http://ocw....>

## **ChemLab - 12. Electrochemistry - Voltaic Cells - YouTube**

1. Given a diagram of a  
simple electrochemical

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cell involving two metal electrodes and the corresponding solution of the metal ions identify: the site of oxidation reduction, the anode, the cathode, movement of electrons, migration of ions, the chemical equation representing the cell reaction.

## **Electrochemical Cells Computer Simulation: Voltaic Cells...**



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### ELECTROCHEMISTRY

**OBJECTIVE:** The objective of the lab was to gain a better understanding of oxidation- reduction reactions, the activity series, and electrochemical cells.

In the lab we compared the electron affinities of different metals, using an electrochemical cell.

### INTRODUCTION:

“Redox” reactions are chemical reactions that

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involve the transfer (loss or gain) of one or more electrons.

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...

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Left: Electrodes:  
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Lead Magnesium Nickel  
Silver Zinc Whodatium  
Pt / Hydrogen. Select  
Solution on Left:  
Solutions: Cadmium  
Nitrate Copper (II)  
Nitrate Iron (II) Nitrate  
Lead (II) Nitrate  
Magnesium Nitrate  
Nickel (II) Nitrate Silver  
Nitrate Zinc Nitrate  
Whodatium (II) Nitrate  
Nitric Acid.

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cell is produced when a  
redox reaction occurs  
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## **Electrochemical Cells Lab Answers 21 - Universitas Semarang**

In this three-part lab,  
these reactionq are

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studied by constructing various electrochemical cells and measuring the voltage generated. From these measurements, a reduction series is generated, the concentration of copper ions in solution determined, and the solubility product of silver chloride calculated.

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