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## **Limiting Reactant Gizmo Answers**

Sometimes during a chemical reaction, one type of reactant will be used up

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before the other reactants. This reactant is the limiting reactant. Using the Limiting Reactants Gizmo™, you can determine which reactant is limiting in various scenarios. To begin, make sure  $H_2 + O_2$  becomes  $H_2O$  is selected.

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The limiting reactant in a chemical reaction is the reactant that A. is used up at the end of the reaction. B. is left over after the reaction. C. has the largest number of atoms. D. has the fewest number of atoms. Explanation: The limiting reactant will be the reactant that is "fully consumed" upon completion of the reaction.

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make sure  $H_2 + O_2$  becomes  $H_2O$  is selected. The small "2" in  $H_2$ ,  $O_2$ , and  $H_2O$  is a subscript. Subscripts represent the number of atoms in a molecule.

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choose the number of molecules of each reactant, and then observe the products created and the reactants left over.

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Limiting Reactants Gizmo™ you can determine which reactant is limiting in various scenarios To begin make sure H<sub>2</sub> O<sub>2</sub> becomes H<sub>2</sub>O is selected The small “2” in H<sub>2</sub> O<sub>2</sub> and H<sub>2</sub>O is a subscript Name Kegan Graves Date

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reactants. This reactant is the limiting reactant. Using the Limiting Reactants Gizmo™, you can determine which reactant is limiting in various scenarios. To begin, make sure H<sub>2</sub>+O<sub>2</sub> becomes H<sub>2</sub>O is selected.

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Unbalanced equation:  $\text{Fe} + \text{HCl} \rightarrow \text{FeCl}_3 + \text{H}_2$   
a) HCl is the limiting reactant, and 3.83 mo... View Answer. Metallic sodium reacts vigorously with liquid bromine. Suppose 1 kg of sodium is brought into...

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limiting reactant = 0.2 excess reactant

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= H<sub>2</sub> 8. 12 mol H<sub>2</sub> (g) 4 mol O<sub>2</sub> (g) 8 mol H<sub>2</sub>O(g) + 4 mol H<sub>2</sub> left over  
limiting reactant = O<sub>2</sub> excess reactant = H<sub>2</sub>

= H<sub>2</sub> 9. 10 mol H<sub>2</sub> (g) 3 mol O<sub>2</sub> (g) 6 mol H<sub>2</sub>O(g) + 4 mol H<sub>2</sub> left over  
limiting reactant = O<sub>2</sub> excess reactant = H<sub>2</sub>

= H<sub>2</sub> 10. 8 mol H<sub>2</sub> (g) 3 mol O<sub>2</sub> (g) 6 mol H<sub>2</sub>O(g) + 2 mol H<sub>2</sub>

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H<sub>2</sub>O is a subscript. Subscripts represent the number of atoms in a molecule. Page

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Group Quiz , Stoichiometry lab due  
Friday: 12.5 Stoichiometry Lab-answer 2

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pre-lab questions

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